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Draw Lewis structures for the following molecules or ions. If the molecule or ion has resonance forms, draw all of them.

1) BCN $16 e^-$

B#C#N

2) CF (ion) $2(1) + 1 = 10 e^-$ This is a CATION so SUBTRACT electrons from the count to make sure we have a +1 charge!

[C+]#F

3) SeF $34 e^-$

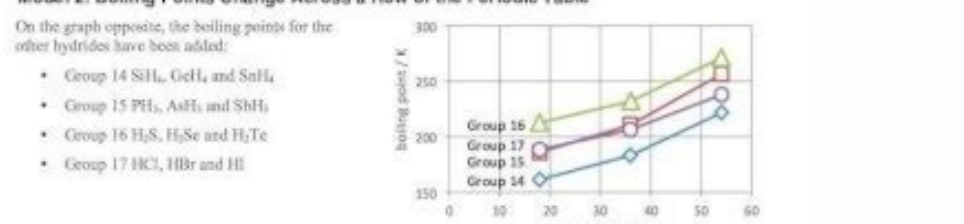
F-Se(F)(F)-F Expanded valence

4) HNO (ion) $14 e^-$

[O-]#N=O ↔ O=N[O-] As shown in your notes with nitrous acid, OXYACIDS have at least one HYDROGEN atom bonded directly to an OXYGEN atom.

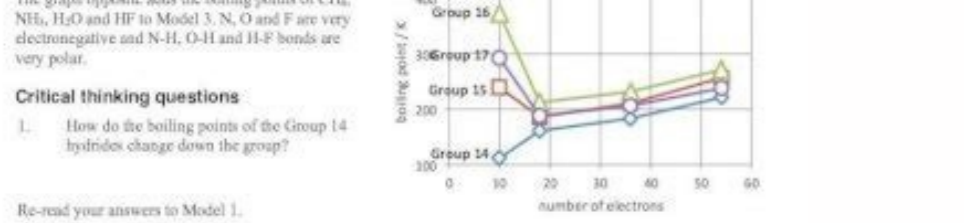
Answers to Lewis structures practice set.

Model 2: Boiling Points Change Across a Row of the Periodic Table



- Critical thinking questions
- Calculate the number of electrons in SiH4, PH3, H2S and HCN.
 - Would you expect a significant change in the strength of dispersion forces in these molecules?
 - What is the molecular shape of PH3, H2S and HCN?
 - Do PH3, H2S and HCN have dipole moments?
 - Why is the boiling point of SiH4 lower than that of PH3, H2S and HCN?
 - Is the boiling point of SeH4 (54 electrons) higher or lower than the boiling point of PH3 (18 electrons)?
 - Explain your answer to Q6, making sure that it is consistent with your answers to Q1 and Model 1.

Model 3: Anomalous Boiling Points of NH3, H2O and HF



- Critical thinking questions
- How do the boiling points of the Group 14 hydrides change down the group?

Answers to Model 2 and Model 3 questions.

Lewis Structures and Molecular Shape

CH4

CH4

NH3

overlap of 2 orbitals

resulting bond (1-1)

Based on their boiling points, which of the following compounds has the largest dipole moment? CS2, CCl4, HI, KI, Cl2, F2, Na2O, H2O, SiO2, CO2, CH4, NH3, H2S, SO2.

Rank the ionic bond strength for the following ionic formulas, 1 being strongest: CS2, CCl4, HI, KI, Cl2, F2, Na2O, H2O, SiO2, CO2, CH4, NH3, H2S, SO2.

Intermolecular forces: Forces that involve covalent or ionic bonding, and weaker interactions that occur between the molecules of a substance. Predict which compound in each pair will have the higher melting point. Improve your knowledge of intermolecular forces with an interactive quiz and printable worksheet.

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